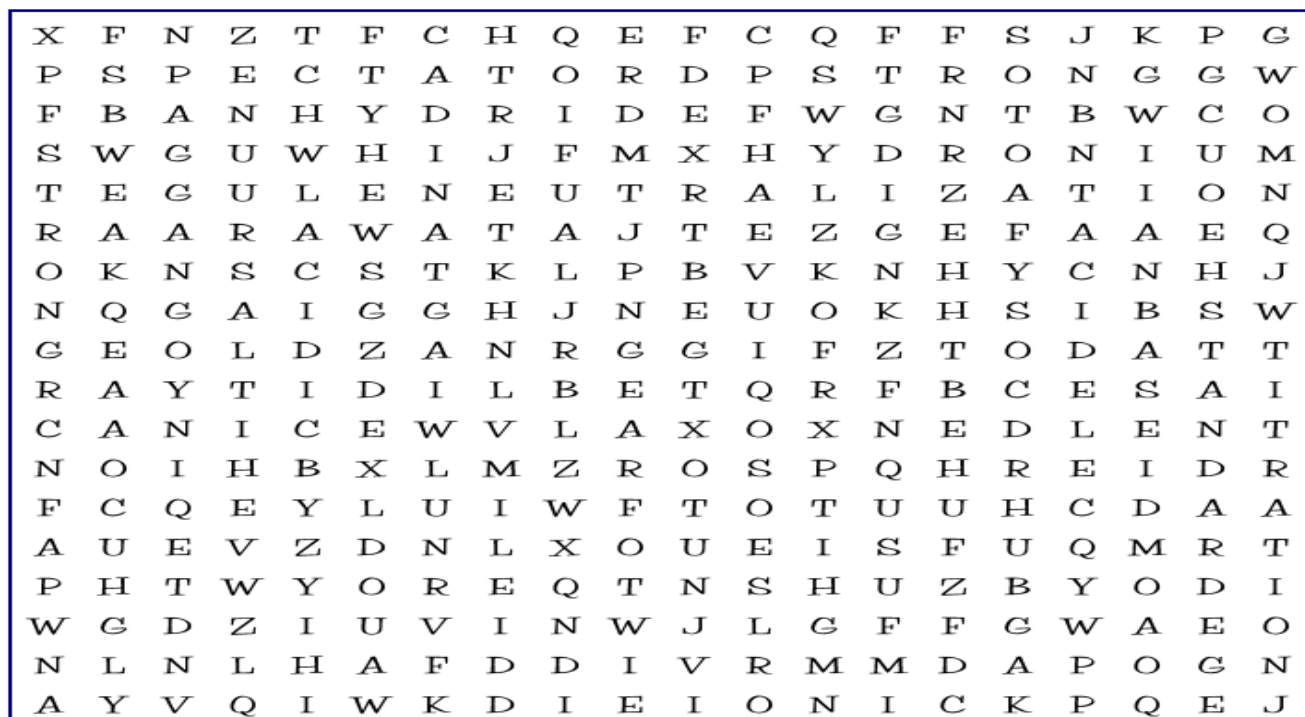


Name: _____

Date: _____

Instructions: Complete the word search puzzle. Use the clues to help you identify the words.**Clues**

- _____ acid - An acid in which almost all the molecules remain as molecules when placed into a water solution.
- Acidic _____ - A nonmetal oxide that reacts with water to form an acid.
- _____ ionic equation - The equation that results when ions common to both sides of the equation are removed, usually from an ionic equation.
- _____ acid - An acid that is completely ionized in water; no molecules exist in the water solution.
- A substance that produces hydroxide ions when it dissolves in water.
- _____ reaction - The reaction of an acid with a base, so called because the properties of both the acid and base are diminished or neutralized.
- _____ solution - A solution of known molarity used in a titration.
- _____ ion - A hydrogen ion attached to a water molecule.
- A substance that produces hydronium ions when dissolved in water.
- The process where ions form from a covalent compound.
- _____ equation - An equation in which substances that primarily exist as ions in solution are shown as ions.
- Basic _____ - A metal oxide that reacts with water to form a base.
- _____ base - A base in which most of the molecules do not react with water to form ions.
- A solution that resists changes in pH when moderate amounts of acids or bases are added to it.
- _____ base - A base that is completely dissociated into separate ions when dissolved in water.
- A mathematical scale in which the concentration of hydronium ions in a solution is expressed as a number from 0 to 14.
- _____ hydrogen - In an acid, any hydrogen that can be transferred to water.
- The process of determining the molarity of an acid or base by using an acid-base reaction where one reactant is of known molarity.
- The general term used in chemistry to describe the ionic compound formed from the negative part of an acid and the positive part of a base.
- _____ ion - An ion that is present in solution but does not participate in the reaction.